umes of BiBr$_3$ (l)[7] and BiCl$_3$ (l)[6] are, significantly smaller than the sum of the molar volumes of their respective elements from which they are formed. On this basis, BiI$_3$ (l) would show a tendency to disproportionate at high pressures whereas BiBr$_3$(l) and BiCl$_3$(l) become more stable. This difference in stability with pressure could account for the difference in behavior of the log $\varkappa$ vs $1/T$ and $\varkappa$ vs. P curves of BiI$_3$ from those of BiBr$_3$. and BiCl$_3$.

The activation energies ($E_\varkappa$) for BiCl$_3$, BiBr$_3$, and BiI$_3$ at $P \leq 0.1$ kbar were estimated from the slope of the log $\varkappa$ vs. $1/T$ curves for these salts from data given by Grantham and Yosim[2a] at temperatures immediately above the melting point. These estimated values of $E_\varkappa$ at $P \leq 0.1$ kbar are compared with $E_\varkappa$ at $P = 5.4$ kbars in Table III. The activation energies at these two pressures are essentially identical in the cases of BiCl$_3$ and BiBr$_3$; for BiI$_3$, $E_\varkappa$ at 5.4 kbars is almost four times its value at $P = 0$. This difference in $E_\varkappa$ between bismuth triiodide and the tribromide and trichloride may be due to a greater effect of pressure upon ionic association in the case of BiI$_3$.